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1. Using carbon-13 and carbon-12, approx mass of neutron =  $13.0034 - 12 = 1.0034$  amu. approx mass of  $^{14}\text{C} = 13.0034 + 1.0034 = 14.0068$  amu.

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nickel (Ni) atoms have in common? ChemActivity 1 The Nuclear Atom 3 . ... electron,  $5.486 \times 10^{-4}$   
amu, to show that the mass of a  $^{12}\text{C}$  atom is less than the sum of the masses of the constituent particles.

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