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of ammonia produced when...

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Answers: Moles and Stoichiometry Practice Problems 1) How many moles of sodium atoms correspond to  $1.56 \times 10^{21}$  atoms of sodium?  $1.56 \times 10^{21}$  atoms Na  $\times$   $1 \text{ mol Na}$

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=  $2.59 \times 10^3$  mol Na  $236.022 \times 10$  atoms Na

2) Determine the mass in grams of each of the following: a. 1.35 mol of Fe  $1.35 \text{ mol Fe} \times 55.845 \text{ g Fe} = 75.4 \text{ g Fe}$  1 mol Fe b. 24.5 mol O

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Response Test: 1. 30.40 g. N X 1 mol

N/14.01g N = 2.17 moles N. 69.60 g O X 1

mol. O /16.00 g O = 4.35 moles O.

4.35/2.17. = 2.00 2.17/2.17 = 1.00 So O is 2X

more present than N. Empirical formula

must be NO. 2. Mass.

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many grams of nitric acid,  $\text{HNO}_3$ , ...

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2) 19. The mass of element X found in 1.00



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mole of each of four different compounds is  
28.0 g, 42.0 g, 56.0 g, and 70 g, ...

Chapter 12 Stoichiometry Test Answer Key  
Practice Test Ch 3 Stoichiometry Name Per.  
Remember it is a MC test, use the answers ...  
Practice Test Ch3 Stoichiometry (page 3 of  
3) 1. d It might be easiest to balance the

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equation with mostly whole numbers:  $2 \text{NH}_3 + \text{O}_2 \rightarrow 2\text{NO} + 3\text{H}_2\text{O}$  ... 7. c First you must realize this is a limiting reactant problem. Found: 3 Jan 2020 | Rating: 89/100

## Stoichiometry Practice Problems Level 1 Answers

20 Then do some stoichiometry using

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“ easy math ” 16 g of methane (MM = 16) is 1 mole and 1 mole of methane will produce 1 mole of  $\text{CO}_2 = 44 \text{ g}$ , and 2 moles of  $\text{H}_2\text{O}$  which is 36 g for a total of 80 g 4. d  
Balance:  $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$   
5. d Balance:  $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$

Practice Test Ch 3 Stoichiometry Name Per

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N2?

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Aluminum bromide can be prepared by the

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reaction of aluminum metal with bromine gas shown by the equation:  $2 \text{Al} + 3 \text{Br}_2 \rightarrow 2 \text{AlBr}_3$  Now suppose that 5.6 mol of aluminum reacts with 4.4 mol of bromine. 1. Calculate the mass of aluminum bromide that can be produced from 5.6 mol of Al. 2.

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## Modern Chemistry

Stoichiometry Practice Test Short Answer:

Aluminum bromide can be prepared by the reaction of aluminum metal with bromine gas shown by the equation:  $2 \text{Al} + 3 \text{Br}_2 \rightarrow 2 \text{AlBr}_3$

Now suppose that 5.6 mol of aluminum reacts with 4.4 mol of bromine. 1. Calculate the mass of aluminum bromide



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