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Molecular weight of the gas = 29 ; $\gamma = 1.4$.

$\frac{1.4 - 1}{1.4 + 1} = \frac{0.4}{2.4} = \frac{1}{6}$; $\frac{1}{6} = 0.1667$

$\left[\frac{1}{6} + 1 \right] : 0.528 \text{ p,} = 430 \text{ K} : 0.523 = 253.44$

kPa Fig. 11121 $\frac{1}{6}, 2 \text{ T}_1 = 7 + 1 = 1.4 + \gamma$ do ' 7',

$= 303 \times 0.3333 = 573.3 \text{ K}$; $R = 9 - \frac{1}{6} \text{ S} ; - \frac{1}{6} \times 3$

$= 0.2357 \text{ kJ/ kg}^\circ\text{C}$; $R = 0.2867 \times 1.4 = 0.4014$

$\frac{1}{6} = 0.1667$; $\frac{1}{6} = 0.1667$; $1.0034 \text{ kJ/ kg}^\circ\text{C}$

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